

# Download Freezing Point Depression Lab Answers

Transcript of Molecular Mass by Freezing Point Depression Lab. Mass 8.00 g of BHT, an empty test tube, and the test tube with the BHT. Heat the water until 90 degrees Celsius after the test tube is submerged. Once the water is at 90 degrees Celsius remove the test tube. Stir the BHT continuously and record the temperature every 20 seconds. CONCLUSIONS. Explain your answer. The molar mass would be lower because if the freezing point was  $0.3^{\circ}$  lower, then there would be a greater change in temperature, which would result in a larger molality and more moles. There would also be a smaller molar mass. Post-lab Analysis. First, we can see that the freezing points are already established. Therefore, we can find  $\Delta T_{fp}$  for BHT + cetyl alcohol and BHT + unknown. This is because  $\Delta T_{fp}$  is the change of the freezing point between BHT and BHT + whichever substance. We use some simple subtraction for this part.